

Закончить схемы возможных реакций, указать условия их протекания:

1. $\text{Cr}_2\text{O}_3 + \text{Al} \rightarrow$
2. $\text{MgO} + \text{H}_2 \rightarrow$
3. $\text{Na}_2\text{O} + \text{CO} \rightarrow$
4. $\text{CaO} + \text{C} \rightarrow$
5. $\text{ZnO} + \text{CO} \rightarrow$
6. $\text{Fe}_3\text{O}_4 + \text{Al} \rightarrow$
7. $\text{K}_2\text{O} + \text{Mg} \rightarrow$
8. $\text{Cu}_2\text{O} + \text{CO} \rightarrow$
9. $\text{WO}_3 + \text{H}_2 \rightarrow$
10. $\text{MgO} + \text{Ca} \rightarrow$
11. $\text{MgO} + \text{C} \rightarrow$
12. $\text{Li} + \text{H}_2 \rightarrow$
13. $\text{Ca} + \text{H}_2 \rightarrow$
14. $\text{Fe} + \text{S} \rightarrow$
15. $\text{Fe} + \text{Cl}_2 \rightarrow$
16. $\text{Be} + \text{C} \rightarrow$
17. $\text{Sr} + \text{N}_2 \rightarrow$
18. $\text{Na} + \text{P} \rightarrow$
19. $\text{Mg} + \text{O}_2 \rightarrow$
20. $\text{Na} + \text{O}_2 \rightarrow$
21. $\text{Fe} + \text{O}_2 \rightarrow$
22. $\text{Al} + \text{C} \rightarrow$
23. $\text{Al} + \text{I}_2 \rightarrow$
24. $\text{K} + \text{H}_2\text{O} \rightarrow$
25. $\text{Mg} + \text{H}_2\text{O} \rightarrow$
26. $\text{Al} + \text{H}_2\text{O} \rightarrow$
27. $\text{Cu} + \text{H}_2\text{O} \rightarrow$
28. $\text{Zn} + \text{H}_2\text{O} \rightarrow$
29. $\text{Ag} + \text{H}_2\text{O} \rightarrow$
30. $\text{Fe} + \text{H}_2\text{O} \rightarrow$
31. $\text{Mg} + \text{HCl}_{(\text{p-p})} \rightarrow$
32. $\text{Ag} + \text{H}_2\text{SO}_{4(\text{p-p})} \rightarrow$
33. $\text{Fe} + \text{HBr}_{(\text{p-p})} \rightarrow$
34. $\text{Cu} + \text{HI}_{(\text{p-p})} \rightarrow$
35. $\text{Be} + \text{H}_2\text{SO}_{4(\text{p-p})} \rightarrow$
36. $\text{Al} + \text{HCOOH}_{(\text{p-p})} \rightarrow$
37. $\text{Cr} + \text{HCl}_{(\text{p-p})} \rightarrow$
38. $\text{Mg} + \text{H}_3\text{PO}_{4(\text{p-p})} \rightarrow$
39. $\text{Zn} + \text{CH}_3\text{COOH}_{(\text{p-p})} \rightarrow$
40. $\text{Fe} + \text{H}_2\text{SiO}_3 \rightarrow$
41. $\text{Al} + \text{H}_2\text{SO}_{4(\text{p-p})} \rightarrow$
42. $\text{Mg} + \text{KOH} + \text{H}_2\text{O} \rightarrow$
43. $\text{Al} + \text{NaOH} + \text{H}_2\text{O} \rightarrow$
44. $\text{Zn} + \text{KOH} + \text{H}_2\text{O} \rightarrow$
45. $\text{Ba} + \text{NaOH} + \text{H}_2\text{O} \rightarrow$
46. $\text{Be} + \text{KOH} + \text{H}_2\text{O} \rightarrow$
47. $\text{Mg} + \text{LiI}_{(\text{p-p})} \rightarrow$
48. $\text{Al} + \text{AgNO}_{3(\text{p-p})} \rightarrow$
49. $\text{Fe} + \text{MgSO}_{4(\text{p-p})} \rightarrow$
50. $\text{Zn} + \text{CuCl}_{2(\text{p-p})} \rightarrow$
51. $\text{Pb} + \text{Zn}(\text{NO}_3)_{2(\text{p-p})} \rightarrow$
52. $\text{Cu} + \text{HgCl}_{2(\text{p-p})} \rightarrow$
53. $\text{Co} + \text{Pb}(\text{NO}_3)_{2(\text{p-p})} \rightarrow$
54. $\text{Cr} + \text{AlCl}_{3(\text{p-p})} \rightarrow$
55. $\text{Na} + \text{Al}_2(\text{SO}_4)_{3(\text{p-p})} \rightarrow$
56. $\text{Ag} + \text{H}_2\text{SO}_{4(\text{конц.})} \rightarrow$
57. $\text{Al} + \text{HNO}_{3(\text{конц.})} \rightarrow$
58. $\text{Hg} + \text{HNO}_{3(\text{конц.})} \rightarrow$
59. $\text{Fe} + \text{H}_2\text{SO}_{4(\text{конц.})} \rightarrow$
60. $\text{Cu} + \text{HNO}_{3(\text{разб.})} \rightarrow$
61. $\text{Au} + \text{H}_2\text{SO}_4 \rightarrow$
62. $\text{Fe} + \text{H}_2\text{SO}_{4(\text{разб.})} \rightarrow$